**CW #- % Composition, Empirical and Molecular Formulas**

**Name Date**

**Period: 1 2 3**

Directions: Solve each problem, SHOW YOUR WORK AND HAVE ANSWERS WITH CORRECT UNITS AND SIGNIFICANT FIGURES!

% Composition – Find the % of each element by mass in the compound.

1. Copper (II) bromide
2. Sodium hydroxide
3. Potassium permanganate
4. Magnesium nitrate
5. Aluminum sulfite

**Calculate the following:**

1. What is the percent composition of lead in PbCl2?
2. What is the percent of sodium in Na2CO3?
3. Determine the percent composition of Na & Cl in NaCl.
4. What is the percent of silver in AgNO3?
5. A chemist breaks down a compound into 5.4 grams of carbon & 7.2 grams of oxygen. What is the percent composition by mass of each element in the new compound?